

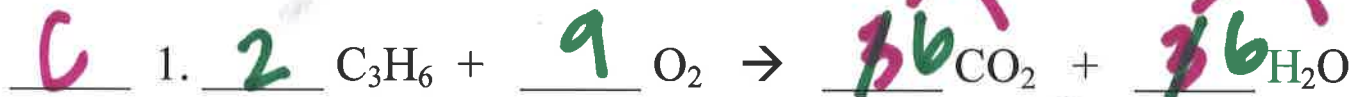
Name: _____

Period: _____

Chapter 11 Exam Review

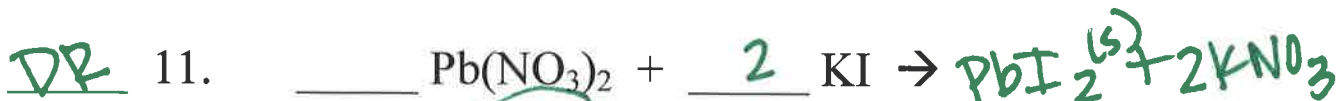
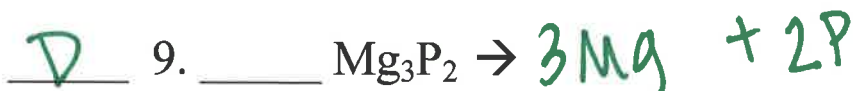
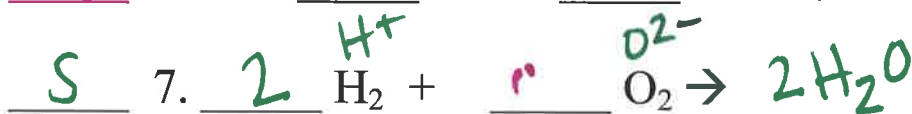
Identify the type of reaction in the first blank and then balance the equations.

Synthesis (S), Decomposition (D), Single-replacement (SR), Double-replacement (DR), Combustion (C).

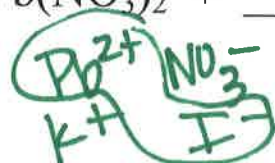


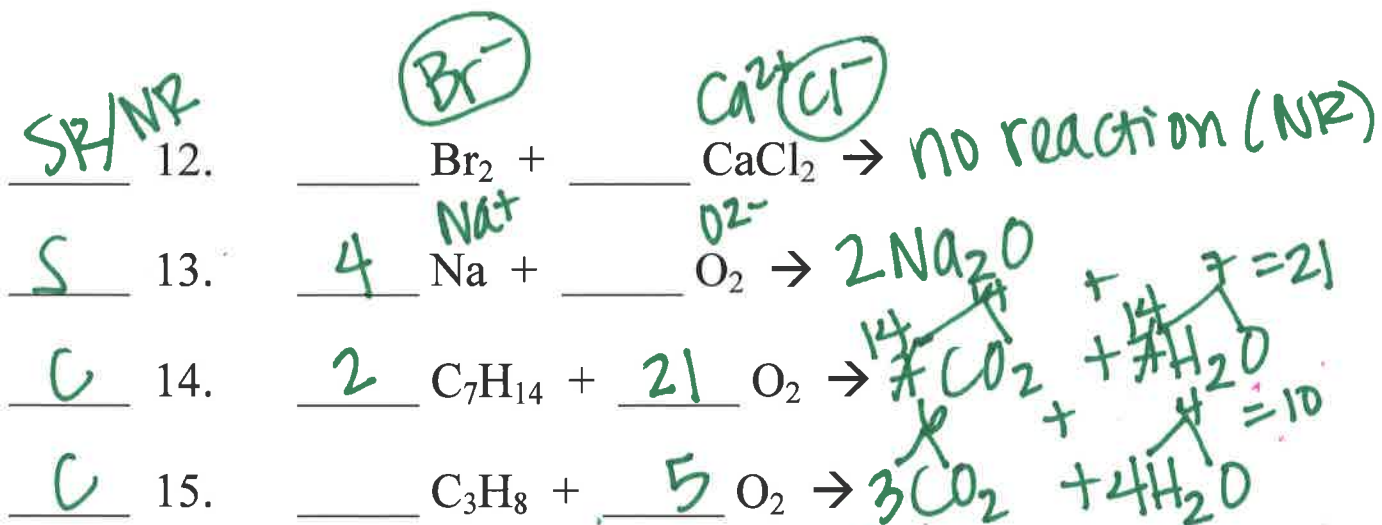
Identify the type of each reaction. If no reaction occurs write N.R.

Complete and balance the following:



diatomics:
H BRONCIE F





Identify the type of reaction. Complete the word equation and write the formula equation for the following reactions. Balance the formula equations.



Balanced equation: $\text{I}_2 + \text{CaCl}_2 \rightarrow \text{NR}$



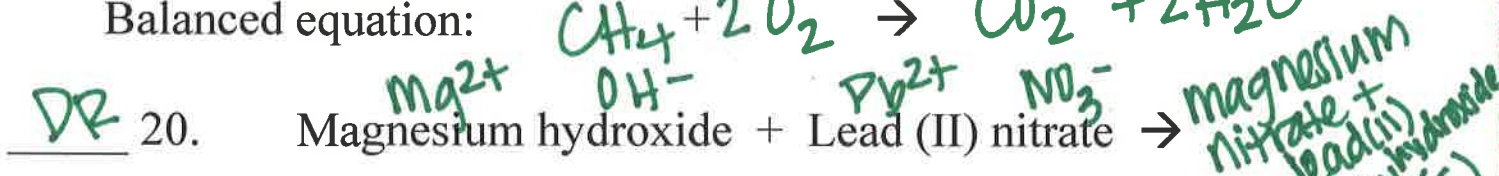
Balanced equation: $2 \text{Al} + \text{Fe}_2\text{O}_3 \rightarrow 2 \text{Fe} + \text{Al}_2\text{O}_3$



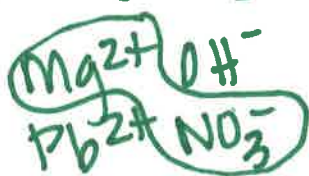
Balanced equation: $3 \text{Sr} + \text{N}_2 \rightarrow \text{Sr}_3\text{N}_2$

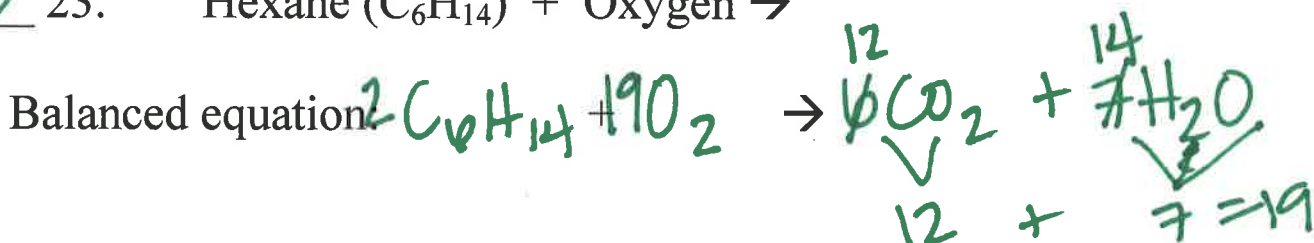
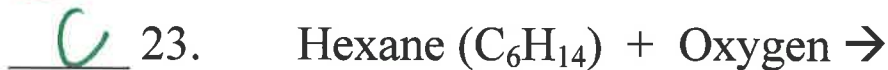
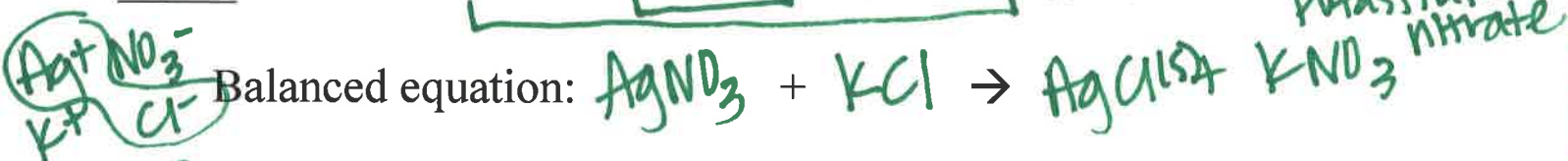
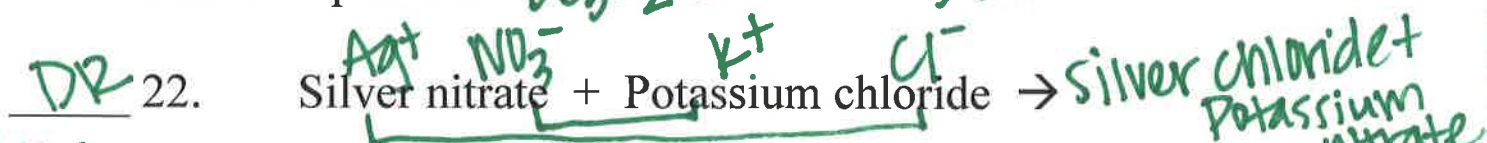
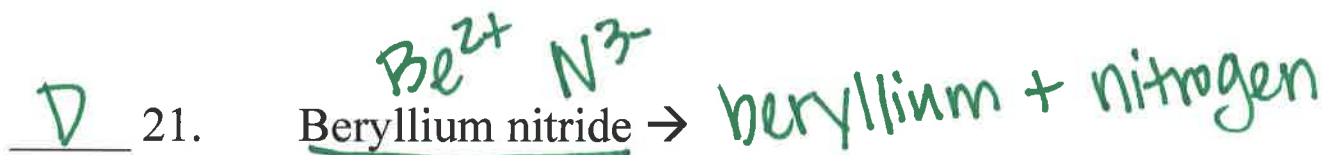


Balanced equation: $\text{CH}_4 + 2 \text{O}_2 \rightarrow \text{CO}_2 + 2 \text{H}_2\text{O}$

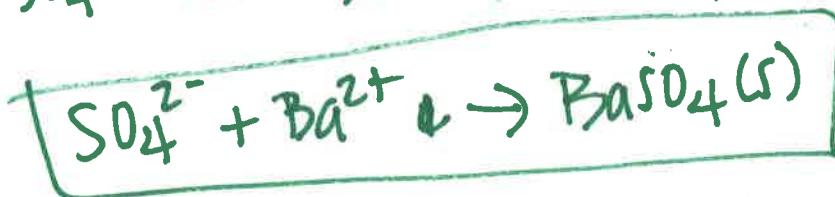
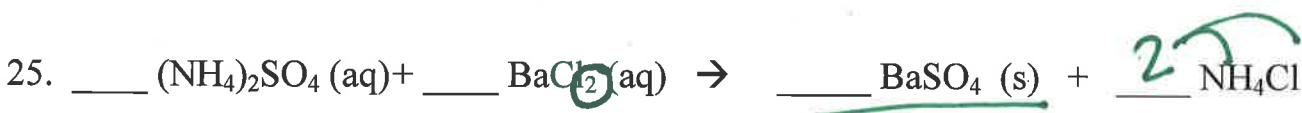
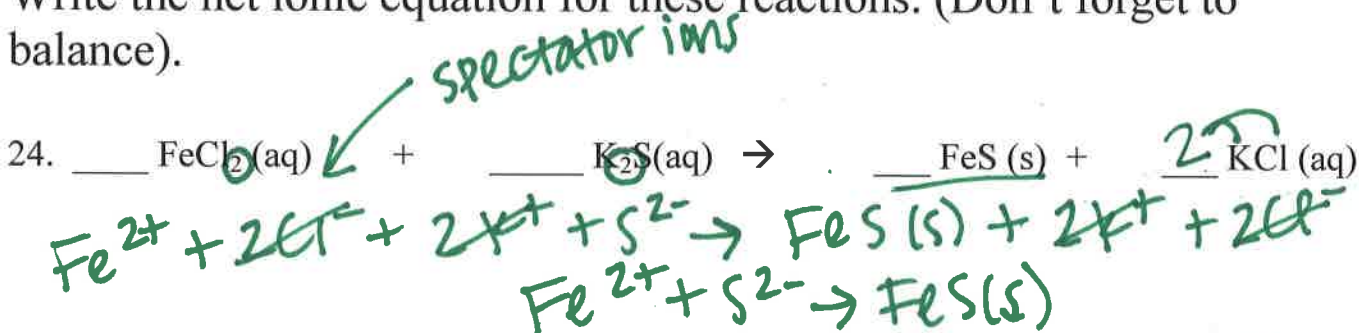


Balanced equation: $\text{Mg}(\text{OH})_2 + \text{Pb}(\text{NO}_3)_2 \rightarrow \text{Mg}(\text{NO}_3)_2 + \text{Pb}(\text{OH})_2(\text{s})$





Write the net ionic equation for these reactions. (Don't forget to balance).



Know what each of these are:

Product $R \rightarrow P$
what is produced in a rxn. (R side of the arrow)

Reactant
what is reacting in a rxn (L side of arrow)

Balanced equation
• follows the law of cons. of mass
• same # & type of atoms on each side of rxn.

Activity series of metals
• used for single replacement rxns.
• the higher it is on the list, the more reactive it is.

Catalyst
• speeds up a rxn
• located on top of arrow (yield) in chem. eqn.

Spectator ions
• cancelled out for overall net ionic eqn.
• not part of overall rxn.

Coefficients
• large #s in front of reactants and/or products

Single-replacement reaction
• an element replaces another element in a compd.
 $A + BC \rightarrow AC + B$ (use activity series)

Combustion reaction
• reactant is O_2 !
 $C_xH_y + O_2 \rightarrow CO_2 + H_2O$

Decomposition reaction
• one compound decomposes into 2 elements or compounds
 $AB \rightarrow A + B$

