

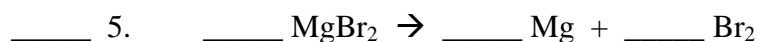
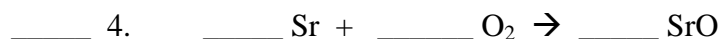
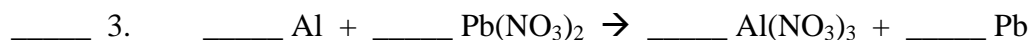
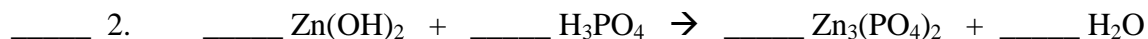
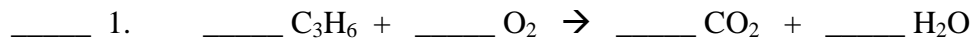
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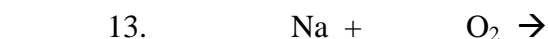
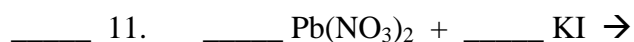
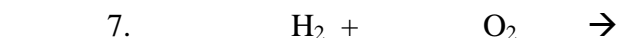
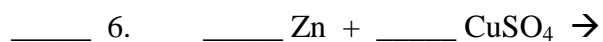
Chapter 11 Exam Review

Identify the type of reaction in the first blank and then balance the equations.

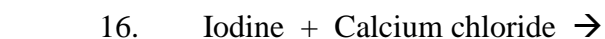
Synthesis (S), Decomposition (D), Single-replacement (SR), Double-replacement (DR), Combustion (C).



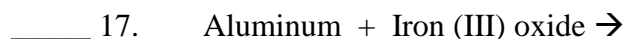
Identify the type of each reaction. If no reaction occurs write N.R. Complete and balance the following:



Identify the type of reaction. Complete the word equation and write the formula equation for the following reactions. Balance the formula equations.



Balanced equation: _____ + _____ \rightarrow



Balanced equation: _____ + _____ \rightarrow

_____ 18. Strontium + Nitrogen →

Balanced equation: _____ + _____ → _____

_____ 19. Methane (CH₄) + Oxygen →

Balanced equation: _____ + _____ → _____

_____ 20. Magnesium hydroxide + Lead (II) nitrate →

Balanced equation: _____ + _____ → _____

_____ 21. Beryllium nitride →

Balanced equation: _____ + _____ → _____

_____ 22. Silver nitrate + Potassium chloride →

Balanced equation: _____ + _____ → _____

_____ 23. Hexane (C₆H₁₄) + Oxygen →

Balanced equation: _____ + _____ → _____

Write the net ionic equation for these reactions. (Don't forget to balance).

24. _____ FeCl₂ (aq) + _____ K₂S(aq) → _____ FeS (s) + _____ KCl (aq)

25. _____ (NH₄)₂SO₄ (aq) + _____ BaCl₂ (aq) → _____ BaSO₄ (s) + _____ NH₄Cl

Know what each of these are:

Product

Balanced equation

Catalyst

Coefficients

Combustion reaction

Reactant

Activity series of metals

Spectator ions

Single-replacement reaction

Decomposition reaction