Homework # 6 Test Review

Chemical Formula	Ionic (I) or Covalent (C)	Chemical Name
1. Sr_3P_2	i	strontium phosphide
2. KI	i	potassium iodide
3. PF_5	c	phosphorous pentaiodide fluoride
4. N_2O	c	dinitrogen monoxide
5. P_4F_{10}	c	tetraphosphorous decafluoride
6. CoI_2	i	cobalt (II) iodide
7. Cr_2O_3	i	chromium (III) oxide
8. $Ba(CN)_2$	i	barium cyanide
9. Al_2S_3	i	aluminum sulfide
10. KOH	i	potassium hydroxide
11. KrF_2	i & c	krypton difluoride
12. $FeCl_3$	i	iron (III) chloride
13. K_2O	& i	potassium oxide
14. Na_2SO_4	i	sodium sulfate
15. $Cu_3(PO_4)_2$	i	copper (II) phosphate
16. $CaBr_2$	i	calcium bromide
17. CI_4	c	carbon tetraiodide
18. $FeCrO_4$	i	iron (II) chromate
19. MgO	i	magnesium oxide
20. $CuCl_2$	i	copper (II) chloride
21. FeO	i	iron (II) oxide
22. $CaSO_3$	i	calcium sulfite
23. $AlPO_4$	i	aluminum phosphate
24. NO_2	c	nitrogen dioxide
25. Na_2O	i	sodium oxide

polyatomic ion

transition metal that needs roman numeral

Chemical Name	Ionic (I) or Covalent (C)	Chemical Formula
1. silver oxide	I	$Ag^+ O^{2-} \Rightarrow Ag_2O$
2. rubidium selenide	I	$Rb^+ Se^{2-} \Rightarrow Rb_2Se$
3. tin (II) chloride	I	$Sn^{2+} Cl^- \Rightarrow SnCl_2$
4. zinc carbonate	I	$Zn^{2+} CO_3^{2-} \Rightarrow ZnCO_3$
5. lead (II) nitrite	I	$Pb^{2+} NO_2^- \Rightarrow Pb(NO_2)_2$
6. nitrogen triiodide	C	NI_3
7. bromine pentafluoride	C	BrF_5
8. tetraarsenic decoxide	C	As_4O_{10}
9. magnesium iodide	I	$Mg^{2+} I^- \Rightarrow MgI_2$
10. gold (I) chloride	I	~~Ag~~ $Au^+ Cl^- \Rightarrow AuCl$
11. strontium fluoride	I	$Sr^{2+} F^- \Rightarrow SrF_2$
12. lithium phosphoride	I	$Li^+ P^{3-} \Rightarrow Li_3P$
13. cesium selenide	I	$Cs^+ Se^{2-} \Rightarrow Cs_2Se$
14. cobalt (II) chloride	I	$Co^{2+} Cl^- \Rightarrow CoCl_2$
15. phosphorous trichloride	C	PCl_3
16. carbon tetrachloride	C	CCl_4
17. magnesium perchlorate	I	$Mg^{2+} ClO_4^- \Rightarrow Mg(ClO_4)_2$
18. lead (IV) iodide	I	$Pb^{4+} I^- \Rightarrow PbI_4$
19. lithium chlorate	I	$Li^+ ClO_3^- \Rightarrow LiClO_3$
20. potassium acetate	I	$K^+ C_2H_3O_2^- \Rightarrow KC_2H_3O_2$
21. tin (IV) oxide	I	$Sn^{4+} O^{2-} \Rightarrow SnO_2$
22. silver sulfide	I	$Ag^+ S^{2-} \Rightarrow Ag_2S$
23. rubidium bromide	I	$Rb^+ Br^- \Rightarrow RbBr$
24. copper (II) nitride	I	$Cu^{2+} N^{3-} \Rightarrow Cu_3N_2$
25. calcium nitrate	I	$Ca^{2+} NO_3^- \Rightarrow Ca(NO_3)_2$

polyatomic ion

prefixes