

## Success 24/7 Chemistry: All Gas Law Practice

1. At 210.0 °C a gas has a volume of 8.00 L. What is the volume of this gas at -23.0 °C?
2. If 10.0 liters of oxygen at STP are heated to 512 °C, what will be the new volume of gas if the pressure is also increased to 1520.0 mmHg?
3. A gas occupies 11.2 liters at 0.860 atm. What is the pressure if the volume becomes 15.0 L?
4. A 32.6 mL sample of hydrogen gas is collected over water using a technique known as water displacement. The temperature of the gas is 295 K. The total pressure of the mixture (hydrogen and water vapor) is 785.2 torr. At 295 K the vapor pressure of water is 19.8 torr. Find the pressure of the hydrogen gas alone.



